

Prebiotic Chemistry: What We Know, What We Don't

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Abstract How life on Earth began remains an unexplained scientific problem. This problem is nuanced in its practical details and the way attempted explanations feedback with questions and developments in other areas of science, including astronomy, biology, and planetary science. Prebiotic chemistry attempts to address this issue theoretically, experimentally, and observationally. The ease of formation of bioorganic compounds under plausible prebiotic conditions suggests that these molecules were present in the primitive terrestrial environment. In addition to synthesis in the Earth's primordial atmosphere and oceans, it is likely that the infall of comets, meteorites, and interplanetary dust particles, as well as submarine hydrothermal vent synthesis, may have contributed to prebiotic organic evolution. The primordial organic soup may have been quite complex, but it did not likely include all of the compounds found in modern organisms. Regardless of their origin, organic compounds would need to be concentrated and complexified by environmental mechanisms. While this review is by no means exhaustive, many of the issues central to the state of the art of prebiotic chemistry are reviewed here.

Keywords Prebiotic chemistry · Origin of life · Organic chemistry · Biomolecules

Introduction

“Prebiotic chemistry” can be understood to mean various things: chemistry which occurred before life began or the chemistry which led to life on Earth, and possibly on other

planets. Workers in the field practically define it as naturally occurring, mainly organic, chemistry in planetary or other solar system environments, which *may* have contributed to the origin of life on Earth, or elsewhere. The terms “abiotic chemistry” (chemistry which takes place in the absence of biology) and “prebiotic chemistry” are in some senses synonymous. Since it is generally assumed that the universe is not goal directed, and since it is not known what processes led to the origin of life, the study of prebiotic chemistry almost certainly includes both productive and nonproductive chemical processes. This review places this chemistry in a historical and cosmic context and details some of the known reactions thought to be important. However, the interested reader is referred to more technical texts (Miller and Orgel 1974; Cleaves 2008; Cleaves and Lazcano 2009) and references therein.

The origin of life remains unexplained despite decades (or perhaps centuries, depending on where one historically marks the starting point) of research. A considerable amount of study has provided compelling details as to how it *might* have occurred and whether it is likely to be a universal phenomenon. The development of modern thought on the topic has a long and winding history and has been modified to adapt to developments in other fields, including astronomy, biology, chemistry, and geology. This review covers the synthesis of small organic compounds; however, it should be borne in mind that the major uncertainties revolve around how these compounds self-organize into self-replicating systems.

Birth of a Model

At the beginning of the twentieth century, several less-than-mainstream scientists believed they were able to demonstrate spontaneous generation (Hanczyc 2008). The majority of scientists viewed their experiments with skepticism, partly

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due to Pasteur's masterful negative demonstration in the nineteenth century (Fry 2000). This left a paradox: life appeared to have a common ancestor, yet the origin of life could not be reproduced experimentally. Various explanations for this paradox were presented; for example Svante Arrhenius proposed that, while it was then known that the Earth was not infinitely old, the universe might be, and it was thus possible that, like the universe, life was infinitely old. The spores of some type of bacterium-like seed organism might constantly inoculate new planets with life, which was then able to evolve into the higher forms observed today (Fry 2000).

An alternative and more experimentally addressable proposal was put forth almost simultaneously by Oparin (1924) and Haldane (1929), suggesting that the early Earth may have been markedly different, providing chemical conditions facilitating spontaneous generation. Especially important were Oparin's suggestions that organic compounds, presumably required for the origin of the first organisms, could have been delivered to the Earth on meteorites or that they may have been generated in the early atmosphere, had it been reducing.

In the 1950s, Nobel prize-winning American chemist Harold Urey proposed that the early atmosphere would have been largely reducing, based on the preponderance of hydrogen in the solar system (Urey 1952). Chemists use the terms “oxidizing” and “reducing” to describe the elements that carbon typically bonds with in an atmosphere. Carbon can form four bonds with other atoms: an oxidizing atmosphere, carbon is bound via two double bonds with O to give CO₂ (an oxidized form of carbon) or, in a reducing atmosphere, with four H atoms to give CH₄ (a reduced form of carbon). The import of this is that it is considerably easier to make organic compounds abiotically from reduced gases than from oxidized ones (see below).

A young graduate student, Stanley Miller, after attending a lecture by Urey on his ideas, suggested to Urey trying an experiment to test this model as a thesis project. The

experiment itself was simple: a flask filled with reduced gases and heated circulating water (simulating the Earth's hydrologic cycle) was subjected to an electrical discharge, simulating lightning (Fig. 1).

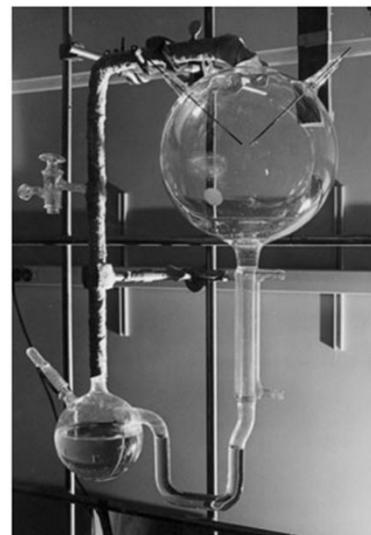
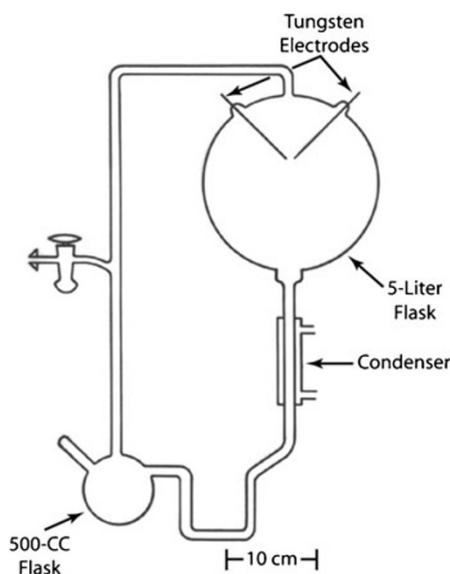
Chemical analysis of the resulting products revealed a surprisingly efficient synthesis of a number of amino acids found in biochemistry, including glycine, and racemic alanine and aspartic acid (Miller 1953). While there had been earlier laboratory demonstrations of organic synthesis using electric discharges (see, for example, Löb 1913), this experiment is generally considered the first conducted in the context of trying to understand the origin of life. It is thus widely deemed to have opened the modern experimental period of research into the mechanism of the origin of life and to have been the first intentional example of “prebiotic chemistry.”

Coincidentally, Watson and Crick published their structure for the DNA double helix within a week of the publication of Miller's results (Watson and Crick 1953). Until that time, it had been widely debated whether proteins or nucleic acids were the carriers of genetic inheritance (though the evidence was strongly in favor of the latter): the structure of DNA left little doubt. This close historical juxtaposition of discoveries reveals a common motif in prebiotic chemistry and in origins-of-life models in general: discoveries in other fields frequently drive advances in origin of life models.

All modern organisms are composed of cells, a fact recognized since the 1800s. However, the understanding of the molecular-scale functioning of cells has undergone remarkable development, allowing more precise questions to be posed regarding the origin of the first living cell.

Modern cells share a variety of common biochemical characteristics which are believed to be evidence of a common ancestry: all are bounded by lipid membranes, all contain DNA, all contain ribosomes, and all use a very similar genetic code to produce the protein enzymes that carry out cellular metabolism.

Fig. 1 The Miller–Urey apparatus. The 500-liter flask contained heated water, evaporating upward and coming into contact with the gases contained in the five-liter flask. The molecules created by contact with the electric discharge were returned to the 500-cubic centimeter flask by the condenser. Reproduced from Lazcano and Bada (2003)



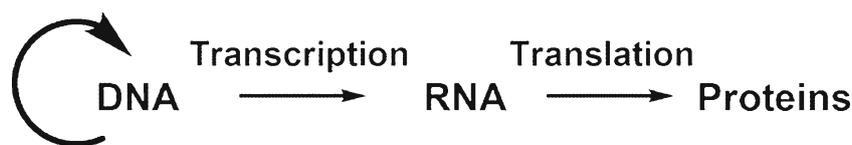


Fig. 2 The central dogma of information flow in biology. The *circular arrow* leading back to DNA represents the fact that DNA is a template for its own replication, a process which is mediated by protein and RNA

The general flow of information in cells is known as the “central dogma” of biology (Fig. 2), which holds that DNA encodes various RNA molecules, which in turn are used to make coded proteins. These RNA molecules include ribosomal RNA (rRNA), which folds into ribosomes (that also contain numerous structurally important peptides) - the protein-manufacturing machines of the cell, transfer RNA (small folded RNA adapter molecules that ensure the precise amino acid coded by a DNA strand gets incorporated into the proper location in a protein) and messenger RNA (into which coded genetically encoded messages of DNA are transcribed before being translated into proteins). DNA is itself copied using various protein enzymes and small RNA primers.

The fact that proteins are needed to make DNA and DNA is needed to make proteins leads to a “chicken or egg paradox”: which logically would have had to arise first? The central role of RNA in this process led some to speculate that it could be the solution to the paradox (Crick 1968; Orgel 1968; Woese 1968), and this led to the notion of an “RNA World” (Gesteland and Atkins 1993), a putative period in which RNA functioned as both catalyst and genetic molecule. Some scientists interpret the RNA World to mean that life began with a self-replicating RNA molecule, while others interpret it to mean that life passed through a period in which RNA was merely extremely important in biochemistry. The perceived simplicity of the first interpretation has driven considerable research into the prebiotic synthesis of RNA.

As recently as the 1990s, the living world was divided into five kingdoms, the animals, plants, fungi, protists, and bacteria (Whittaker 1969). Comparison of rRNA sequences (which presumably are resistant to evolutionary drift) has revised this classification scheme into three domains, phylogenetically organizable into an evolutionary tree of life (Fig. 3): the Eukarya (comprising the eukaryotes), the Bacteria, and the Archaeobacteria (Woese et al. 1990). This last group has a number of unique characteristics, including specialized forms of metabolism and membrane lipids.

Some analyses of this tree suggest hyperthermophilic archaeobacteria are the oldest organisms on Earth (Wang et al. 2007), which has been used to argue that the types of environments these organisms inhabit presently were the earliest environments for life and thus likely sites of the origin of life. This idea remains controversial (Arrhenius et al. 1999; Gupta 2000). The reconstructed tree suggests that the Last Universal Common Ancestor (LUCA) of all modern biology was a single-celled prokaryote, albeit one with

an already sophisticated and very modern biochemistry, suggesting a prior protracted period of biochemical evolution.

To address the question of the origin of LUCA, it is necessary to examine models for how the Earth formed, what geochemical environments may have been available on the early Earth, and what types of organics could have been produced in each.

History of the Earth and Solar System

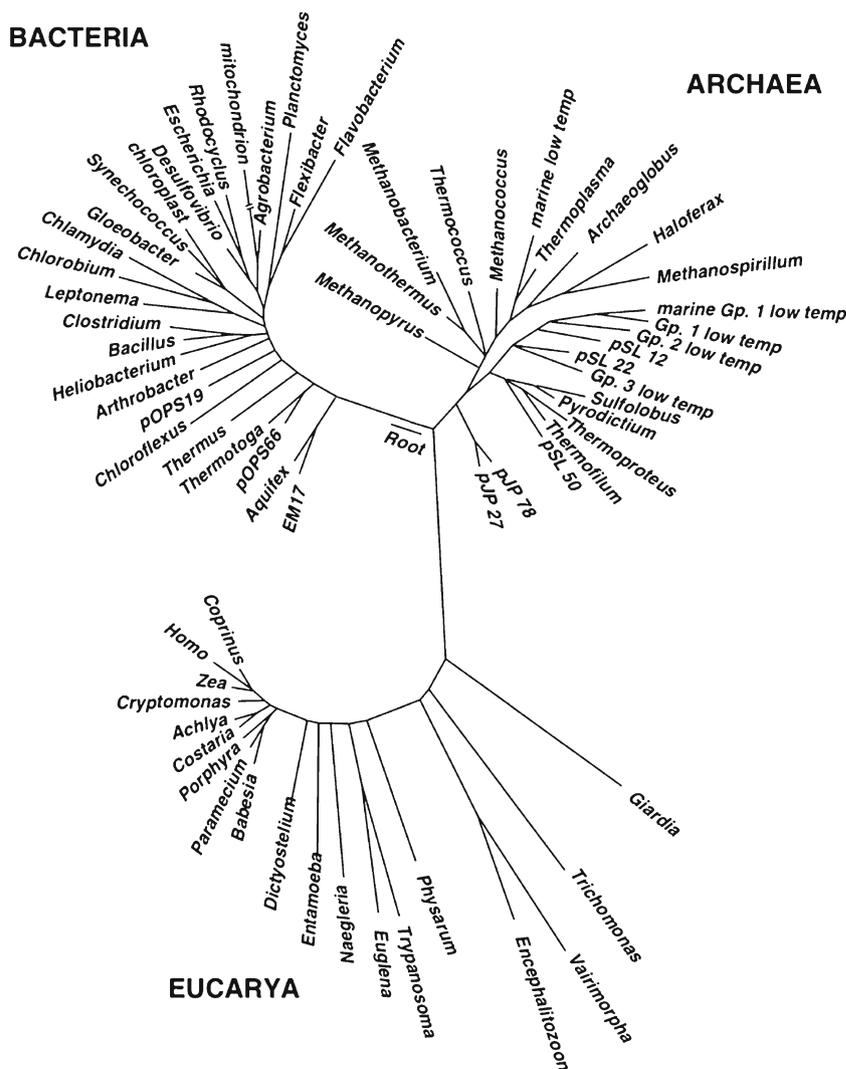
The Earth is believed to be ~4.5 billion years old, only 10–15 million years younger than the solar system itself (Dalrymple 1994). The accepted model for the origin of the solar system is the nebular hypothesis (Montmerle et al. 2006), which holds that the solar system condensed into a disk from a presolar nebula, a cloud of gas and dust composed of remnants of a previous supernova. Due to inelastic collisions during gravitational accretion of small particles, larger bodies formed, gradually coming to occupy orbits at varying distances from the nascent central star, our Sun.

These bodies were dramatically altered in composition when the Sun accreted enough mass to trigger the ignition of thermonuclear fusion in its interior, releasing vast amounts of energetic electromagnetic radiation. Lower boiling point compounds in the surrounding disk were driven outward from the Sun, as in a distillation, each condensing beyond its relative condensation point. Thus, volatile components such as methane, water, and N_2 coalesced into outer solar system bodies such as the gas and ice giant planets (Saturn, Jupiter, Neptune, and Uranus) and the Oort cloud and Kuiper belt comets, while higher boiling-point compounds such as metal oxides and silicates, which make up the bulk of the inner rocky planets, condensed nearer in.

The smaller resultant bodies, comets, meteors, and interstellar dust particles (IDPs) continue to impact the planets presently, and it is likely that this flux was higher around the time that life formed on Earth. The presence of extraterrestrial organic molecules in meteorites, comets, and IDPs is firmly established and has led to proposals that these were sources of organic compounds possibly necessary for the origin of life (Oró 1961; Anders 1989; Chyba et al. 1990; Chyba and Sagan 1992).

The flux of extraterrestrial organics to the early Earth has been estimated based on the lunar cratering record (Chyba and Sagan 1992). These may have contributed significantly to the primitive Earth's prebiotic organic inventory, even if the early Earth's atmosphere were oxidizing or neutral

Fig. 3 The reconstructed Tree of Life based on ribosomal RNA sequences. Figure reproduced from Pace (1997)



(Chyba and Sagan 1992; Thomas et al. 1996). Their components resemble the products of atmospheric synthesis under reducing conditions (Wolman et al. 1972); thus, their compositions are worth examining.

Comets

Comets are mixtures of dust and ice accreted early in the history of the solar system (Festou et al. 2004). In addition to water ice and various inorganic components, the volatile organic components of several comets have been measured spectroscopically (Ehrenfreund and Charnley 2000). Highly reactive organic compounds such as hydrogen cyanide (HCN) and formaldehyde (HCHO), among others, though variable from comet to comet, are often observed in high abundance (Table 1).

Measurement of the hydrogen isotope ratios of cometary water suggests that some of the Earth's oceans may be derived from comets (Chyba 1990; Meier et al. 1998). If this is true, comets could also have delivered organics,

though the survival of these would depend on the nature of the delivery process (Oró 1961; Oró et al. 1980). Assuming that cometary nuclei have a one-gram per cubic centimeter density, a one-kilometer-diameter comet would

Table 1 The abundance of small molecules relative to water in comet Hale–Bopp as measured spectroscopically

Compound	Abundance
H ₂ O	100
CO	23
CO ₂	6
CH ₃ OH	2.4
H ₂ CO	1.1
NH ₃	0.7
CH ₄	0.6
HCN	0.25
HCONH ₂	0.02
CH ₃ CHO	0.02

Adapted from Bockelée-Morvan et al. (2000)

contain 2×10^{11} moles of HCN, or 40 nanomoles per square centimeter of the Earth's surface. This is comparable to the yearly production of HCN in a reducing atmosphere from electric discharges and would be important if the Earth did not have a reducing atmosphere, assuming complete survival of the HCN on impact.

While comets are extremely cold, it appears that some aqueous-phase organic reactions have occurred in them, as evidenced by the detection of the simplest amino acid, glycine, in particles returned from comets (Elsila et al. 2009); thus comets could have delivered some more complex compounds as well.

Meteorites

Meteorites represent generally less volatile remnants of the early solar system, i.e., objects which formed closer to the Sun and underwent more significant thermal processing (Lauretta and McSween 2006). Their compositions range from metallic to stony. The latter category includes a class with which prebiotic chemists are particularly fascinated, the carbonaceous chondrites (CCs), which contain a significant organic component, usually one to two % by mass (Alexander et al. 2007). Besides a few recovered cometary grains, CCs remain the best-studied bona fide examples of “prebiotic chemistry,” represented by several hundred exemplars in various curated collections around the world. CC organic material is variable in composition, but typically 70–99% is a complex high molecular weight kerogen-like material, with the remainder being small soluble organic compounds (Pizzarello et al. 2006).

A variety of organic compounds have been identified in CCs, including many found in modern biochemistry (Pizzarello et al. 2006). That these compounds are indigenous to the meteorites, and not terrestrial contamination, is suggested by the facts that they have unusual isotopic ratios and include types of compounds not typically found in biochemistry; and that compounds with stereocenters are found in nearly equal quantities with respect to their optical isomers, with some notable exceptions (Pizzarello and Cronin 2000; Glavin and Dworkin 2009). A brief summary of the types and relative abundances of compounds identified to date is shown in Table 2.

IDPs

The input from IDPs may have been more important than that of comets or meteorites. Their present infall rate is large, and on the primitive Earth, it may have been greater by a factor of 100 to 1,000 (Love and Brownlee 1993). The organic composition of IDPs is poorly understood (Maurette 1998); the only molecules that have been identified to date are polycyclic aromatic hydrocarbons and α -aminoisobutyric acid (Gibson

Table 2 Organic compounds detected in the Murchison carbonaceous chondrite

Class	Concentration (parts per million)	Compounds identified
Aliphatic hydrocarbons	>35	140
Aromatic hydrocarbons	22	87
Polar hydrocarbons	<120	10
Carboxylic acids	>300	48
Amino acids	60	74
Hydroxy acids	15	38
Dicarboxylic acids	>30	44
Dicarboximides	>50	2
Pyridine carboxylic acids	>7	7
Sulfonic acids	67	4
Phosphonic acids	2	4
N-Heterocycles	7	31
Amines	13	20
Amides	Nd	27
Polyols	30	19
Imino acids	nd	10
Total	778	565

Adapted from (Pizzarello et al. 2006)

nd not determined

1992; Clemett et al. 1993). Heterogeneous organic polymers loosely termed tholins (also produced by electric discharges, ionizing radiation, and ultraviolet light) could be major components of IDPs. Amino acids are released from tholins on acid hydrolysis (Khare et al. 1986); tholins could thus also be a source of organics. On entry to the Earth's atmosphere, IDPs could be heated and their tholins pyrolyzed, creating HCN and other molecules, which could then participate in terrestrial reactions (Mukhin et al. 1989; Chyba et al. 1990).

Formation of the Earth and the Origins of the Atmosphere and Oceans

As it is widely believed that life requires an aqueous environment, and it is a cornerstone of many ideas regarding the origins of Earth's early organics that the early atmosphere was reducing, it is worth discussing the origin of the Earth's oceans, atmosphere, and crust, as these have a complex interplay which affects their ability to produce organic compounds.

During solar system formation, as asteroids accreted, the heat generated from the radioactive decay of elements (such as ^{26}Al and ^{40}K) was trapped in the interiors of these bodies, which began to warm up. As asteroidal component melted and liquefied, they migrated inward or outward within these bodies depending on their densities. As asteroids accreted into planetesimal-sized objects, the internally trapped heat

became great enough to melt rock and metal, and denser materials migrated inward to form metallic cores, as in the rocky planets.

Since much of this material was reduced iron, the migration of this metal to the Earth's core took with it a great deal of the planet's reducing equivalents. Much of the lighter material, including elements such as C, N, and H, migrated toward the surface, with the oxidation state of these elements determined by the equilibration conditions they were exposed to during migration.

Whether the early atmosphere was ever reducing remains contentious (Tian et al. 2005), but it seems unlikely that it was for very long. An N_2/CO_2 -dominated atmosphere may be the most stable state in the absence of biology. O_2 in the present atmosphere is almost entirely generated from biological photosynthesis. In the absence of biology, the lifetime of O_2 would be extremely short due to its reaction with elements such as iron in the crust (for example, $1.5 O_2 + 2 Fe^{2+} \rightarrow Fe_2O_3$). A consequence of the lack of O_2 in the early atmosphere is that there would have also been little UV-absorbing ozone (O_3), which would have allowed highly energetic bond-breaking UV radiation to reach the Earth's surface (Cleaves and Miller 1998). It is generally believed, in addition, that the early Sun produced a far larger amount of radiation in the UV region (Kasting and Siefert 2002) than at present.

The oxidation state of the early mantle likely governed the distribution of outgassed species. Holland (1962) proposed, based on the Earth accreting homogeneously and cold that the Earth's atmosphere went through two stages: an early reduced stage before differentiation of the mantle and a later neutral/oxidized stage after differentiation. During the first stage, the redox state of the mantle was governed by the Fe^0/Fe^{2+} redox pair. The atmosphere in this stage would be composed of H_2O , H_2 , CO , and N_2 , with approximately $0.27\text{--}2.7 \times 10^{-5}$ atmospheres of H_2 . Once Fe^0 had segregated into the core, the redox state of magmas would have been controlled by the Fe^{2+}/Fe^{3+} pair or fayalite–magnetite–quartz buffer.

If the core differentiated rapidly (via rapid sinking of Fe^0 into the core), the early atmosphere may have resembled the composition of modern volcanic gases (Rubey 1951). Rubey estimated that a CH_4 -dominated atmosphere could not have persisted for more than $10^5\text{--}10^8$ years due to photolysis. The Urey/Oparin atmosphere (CH_4 , NH_3 , H_2O) model is thus based on astrophysical and cosmochemical constraints, while Rubey's model is based on extrapolation of the geological record. Although early theoretical work had an influence on research, modern thinking on the origin and evolution of the solar system, the Earth, and its atmosphere and oceans has not been shaped largely with the origin of life in mind. Rather, current origin-of-life theories are usually modified to fit frequently-changing geochemical models.

Light atmospheric gases such as H_2 would have been prone to rapid escape to space due to their low escape velocities, while others such as NH_3 would be rapidly decomposed in the atmosphere by UV photolysis (Ferris and Nicodem 1972; Kuhn and Atreya 1979). Nevertheless, low levels of these compounds could have been maintained at steady state, and significant amounts of NH_3 could have dissolved in the oceans if the pH of the early oceans was lower than the pK_a of NH_3 (~ 9.2 at 25 degrees Celsius) (Bada and Miller 1968).

Water outgassed as steam and slowly condensed as the crust cooled, forming the Earth's surface waters. There is some evidence from zircons (weathering-resistant minerals found in crustal rocks) that liquid water was present as early as 4.4 billion years ago (Valley et al. 2002). The temperature and pH of the early oceans remain poorly constrained, with possible ranges between zero and 100 degrees Celsius and a pH range of 5 to 11 (Kempe and Degens 1985; Morse and Mackenzie 1998).

If a reducing atmosphere was required for terrestrial prebiotic organic synthesis, the crucial question is the source of H_2 . Miller and Orgel (1974) estimated the pH_2 as 10^{-4} to 10^{-2} atmospheres, depending on the various sources and sinks. H_2 could have been supplied to the primitive atmosphere by several sources (Tian et al. 2005), and it is unclear what would have governed this balance.

The oxidation state of the atmosphere is important for the production of HCN, an important reactant in the prebiotic synthesis of purines and amino acids (see below). In CH_4/N_2 atmospheres, HCN is produced abundantly (Chameides and Walker 1981; Stribling and Miller 1987), but in CO_2/N_2 atmospheres, most of the N atoms produced by splitting N_2 recombine with O atoms to form NO_x species (Chameides and Walker 1981). Reduced gas mixtures are generally more conducive to organic synthesis than oxidizing or neutral gas mixtures. Even mildly reducing gas mixtures produce copious amounts of organic compounds, and it seems likely that energy was the not the limiting factor (Stribling and Miller 1987).

Evidence for Life on Earth

Evidence for life on Earth was once restricted to animals leaving visible fossils, which extended back as far as the beginning of the Cambrian period (~ 550 million years ago). Micropaleontologists began to find evidence for Precambrian life in the 1960s in the form of fossilized bacteria in cherts. Presently, the oldest commonly agreed upon fossil microorganisms (it cannot be ascertained whether these organisms were archaeobacteria, bacteria, or perhaps members of another no-longer-existent line of organism) are dated back to approximately 3.45 billion years ago (Schopf 1993). There is evidence in some of the oldest known rock

strata for isotopically light carbon, possibly formed via biological activity, as far back as 3.85 billion years ago (Mojzsis et al. 1996). This notion remains controversial, and there is evidence that nonbiological processes could also produce this signal (Brasier et al. 2002).

Rocks returned from the moon suggest that there was a “late heavy bombardment period” in which there was an anomalously high flux of large meteors hitting the Earth ~3.9 billion years ago (Abramov and Mojzsis 2009), which could have been planet-sterilizing. Thus, using the most conservative evidence for life on Earth and the most generous evidence for liquid water, there is about a one-giga year period of time for life to originate. Using the most controversial evidence for biologically fixed carbon and accepting the late heavy bombardment period as having been planet-sterilizing, there is a mere 50 million years available for the origin of life. Unfortunately, we cannot presently say whether the origin of life requires a week, a month, or a year, much less 50 million or a billion years, highlighting our ignorance of some of the key step in this process (Lazcano and Miller 1994).

Top–Down and Bottom–Up Approaches

Prebiotic chemistry attempts to not only produce organic compounds which could have been used to assemble the first living organisms, but also to explain the self-assembly of the first living organisms. For many researchers, the goal of prebiotic chemistry is the synthesis of a simple living system with some of the common attributes of modern cells, i.e., a lipid membrane, RNA and/or DNA, and small protein-like peptides. For the ~60 years since Miller's pioneering experiment, the synthesis of the components of modern cells has been the somewhat less ambitious goal. While the list of compounds which can be synthesized in the lab is impressive, not all modern cell constituents have been proven to be synthesizable under plausible prebiotic conditions, nor have all of them been found in meteorites. One possible explanation for this discrepancy is that some modern components are products of evolved biochemistry and were not added to the biochemical inventory until well after organisms had developed a considerable degree of complexity (Cleaves 2010).

We do not know which compounds were required for the origin of life. Prebiotic chemists tend to focus on compounds which are present in modern biochemistry, ignoring the large fraction of compounds found in meteorites or produced in simulations that are not found in biology. A recent study of the Murchison CC using sophisticated analytical instruments revealed the presence of many as around 14 million distinct low molecular weight organic compounds (Schmitt-Kopplin et al. 2010). This can be contrasted with the approximately 1,500 common metabolites found in contemporary cells (Morowitz 1979) and the some

600 small molecules positively identified in the Murchison meteorite to date (Table 2).

Our ignorance regarding the nature of the compounds required for the origin of life, though, does not stop us from attempting to understand how organic compounds form in prebiotic environments.

Prebiotic Syntheses of Biochemicals

Amino Acids

A variety of prebiotic processes can form amino acids, for example, Miller–Urey-type electric discharge experiments (Miller 1953) and reactions of HCN in water (Ferris et al. 1974), among others, and amino acids are found in a variety of CCs (Martins et al. 2007). One of the likely principal mechanisms of formation of amino acids in these samples is the Strecker synthesis (Fig. 4a), named for Adolf Strecker, a nineteenth-century chemist who was the first to artificially synthesize an amino acid in the laboratory. Evidence for this mechanism in the Miller–Urey (MU) experiment comes from measurements of the concentrations of HCN, aldehydes, and ketones in the water flask produced during the course of the reaction (Fig. 4B), which are derived from the CH₄, NH₃, and H₂ originally introduced into the apparatus. This suggests that amino acids are not formed directly in the electric discharge but are the result of synthesis involving aqueous-phase reactions (Miller 1955).

Both amino and hydroxy acids can be synthesized at high dilutions of HCN and aldehydes in a simulated primitive ocean (Miller 1957). Reaction rates depend on temperature, pH, and HCN, NH₃, and aldehyde concentrations but are rapid on geologic time scales. The half-lives for the hydrolysis of the amino- and hydroxy-nitrile intermediates (the rate-limiting steps in these reactions) are less than 1,000 years at zero degrees Celsius (Miller 1998). Corroborating this notion of rapid synthesis, the amino acids found in the Murchison meteorite were apparently formed in less than 10⁵ years (Peltzer et al. 1984).

The Strecker amino acid synthesis requires the presence of NH₃. As mentioned above, gaseous NH₃ is rapidly decomposed by ultraviolet light, and during Archean times, the absence of a significant ozone layer would have limited NH₃'s atmospheric concentration. However, NH₃ is extremely water soluble (depending on pH) and similar in size to K⁺; thus, it easily enters exchange sites on clays. Thus, a considerable amount of NH₃ may have been dissolved or adsorbed on submerged mineral surfaces.

Spark discharge experiments using CH₄, CO, or CO₂ as a carbon source with various amounts of H₂ show that methane is the best source of amino acids, but CO and CO₂ are almost as good if a high H₂/C ratio is used. Without added H₂, amino acid yields are quite low, especially when CO₂ is the sole carbon source (Stribling and Miller 1987). Recent

and ribose and uracil in RNA. Although some now doubt that RNA itself is prebiotic, numerous laboratory experiments show the ease of formation of purines, pyrimidines, and sugars, albeit in low yield.

Purines

The first evidence that purines could be synthesized abiotically was provided in 1961 when Oró reported the formation of adenine (formally a pentamer of HCN, $C_5H_5N_5$) from concentrated solutions of NH_4CN refluxed for a few days. Adenine was produced up to 0.5% yield along with 4-aminoimidazole-5-carboxamide (AICA) and an intractable polymer (Oró and Kimball 1961). It is surprising that a synthesis requiring at least

five steps should produce such high yields of adenine. The initial step is the dimerization of HCN followed by further reaction to give HCN trimer and HCN tetramer, diaminomaleonitrile (DAMN) (Fig. 6).

Ferris and Orgel (1966) demonstrated that a photochemical rearrangement of DAMN proceeds in high yield in sunlight to give aminoimidazole carbonitrile (AICN) (Fig. 6). Other purines, including guanine, can be produced by variations of this synthesis from AICN and its amide (AICA) and other small molecules generated from HCN (Sanchez et al. 1967, 1968) (Fig. 6). These mechanisms are likely an oversimplification. In dilute solution, adenine synthesis may also involve the formation and rearrangement of precursors such as 2- and 8-cyano adenine (Voet and Schwartz 1983).

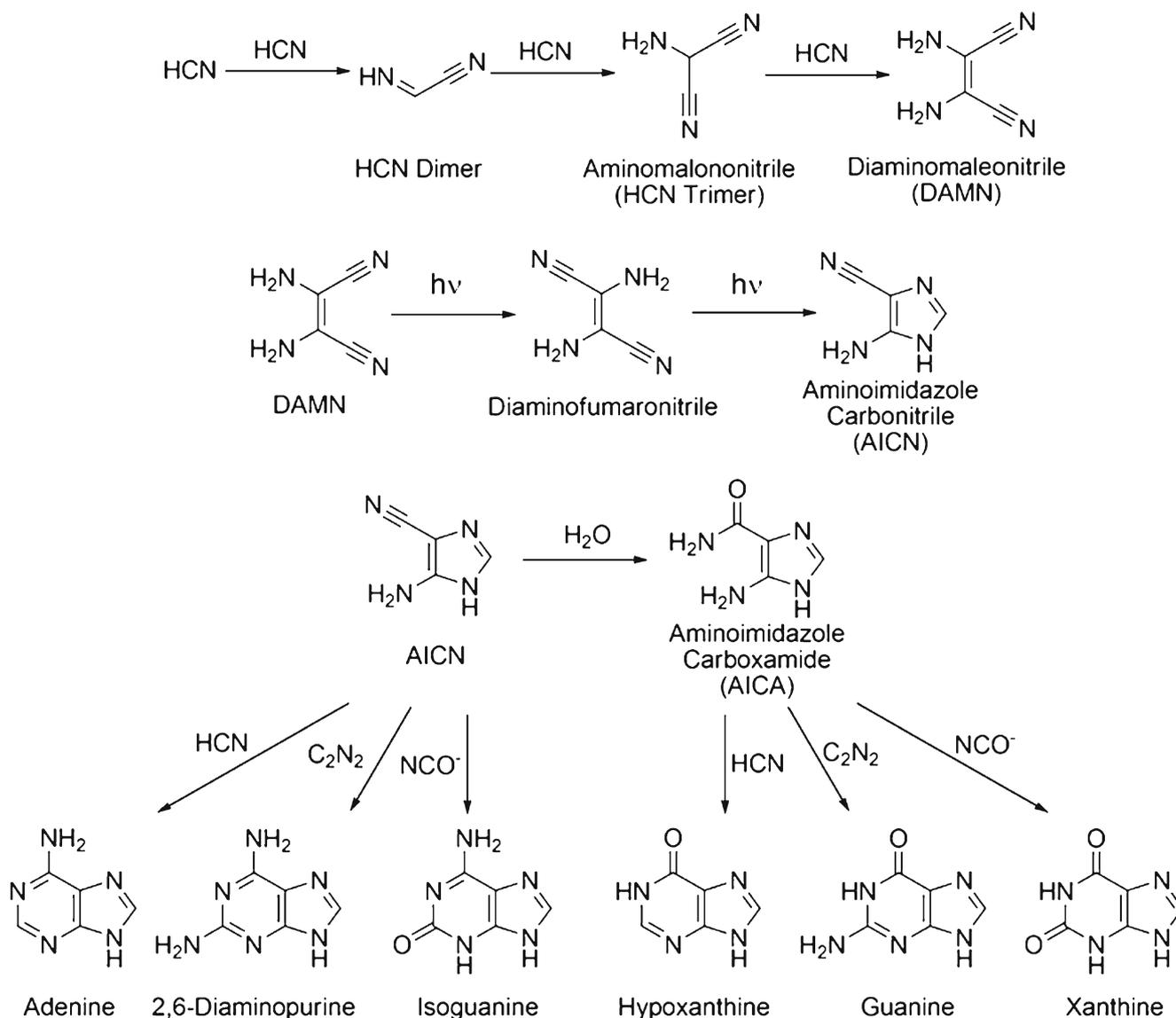


Fig. 6 Proposed mechanisms for formation of DAMN, AICN from DAMN, and purines from the reaction of AICN with small molecules produced in HCN oligomerization and MU-type experiments

The steady-state concentration of HCN in primitive terrestrial waters would have depended on the pH and temperature of the oceans and the input rate of HCN from atmospheric synthesis. Assuming favorable HCN production rates, steady-state concentrations of HCN of 2×10^{-6} molar at pH eight and zero degree Celsius in the primitive oceans have been estimated (Miyakawa et al. 2002b). At 100 degrees Celsius and pH eight, this was estimated at 7×10^{-13} molar. Oligomerization and hydrolysis compete at approximately 10^{-2} molar concentrations of HCN at pH nine (Sanchez et al. 1966a, b), although it has been shown that adenine is still produced from solutions as dilute as 10^{-3} molar (Miyakawa et al. 2002a, b). If the concentration of HCN were as low as estimated, it is possible that DAMN formation may have only occurred on the primitive Earth in eutectic solutions of HCN–H₂O, which would require that some regions of the Earth were frozen. High yields of DAMN are obtained by cooling dilute HCN solutions to negative ten to –30 degrees Celsius for a few months (Sanchez et al. 1966a, b). Production of adenine by HCN polymerization is accelerated by the presence of HCHO and other aldehydes, which could have also been available in the prebiotic environment (Schwartz and Goverde 1982).

The polymerization of concentrated NH₄CN solutions also produces guanine between –80 and –20 degrees Celsius (Levy et al. 1999). Adenine, guanine, and amino acids have also been detected in dilute solutions of NH₄CN kept frozen for 25 years at –20 and –78 degrees Celsius, as well as in the aqueous products of MU experiments frozen for five years at –20 degrees Celsius (Levy et al. 2000).

In addition to producing the biological purines, HCN oligomerization also produces nonbiological purines such as 2,6-diamino- and dioxipurines and the parent compound purine. This same suite of purines has been identified in CCs (Callahan et al. 2011), making a compelling case that similar mechanisms are responsible for their syntheses in CCs.

Formamide (HCONH₂), the hydrolysis product of HCN, has also been shown to produce purines, albeit under more extreme conditions (Bredereck et al. 1961; Saladino et al. 2007).

Pyrimidines

The first “prebiotic” synthesis of pyrimidines investigated was that of uracil (U) from malic acid and urea (Fox and Harada 1961). The prebiotic synthesis of cytosine (C) from cyanoacetylene (HCCCN) and cyanate (NCO[–]) was later described (Sanchez et al. 1966a, b; Ferris et al. 1968) (Fig. 7). HCCCN is produced by the action of spark discharges on CH₄/N₂ mixtures, and NCO[–] is produced from cyanogen or the decomposition of urea.

The high concentrations of NCO[–] required for this reaction may be unrealistic, as it rapidly hydrolyzes to CO₂ and NH₃. Urea is more stable, depending on the concentrations of NCO[–] and NH₃. It was later found that the reaction of dilute cyanoacetaldehyde (CAA) (produced from the hydrolysis of HCCCN) with urea concentrated by evaporation in laboratory simulations of drying beaches gives large (>50%) yields of C (Robertson 1995) (Fig. 7). Evaporating CAA with guanidine produces 2, 4-diaminopyrimidine in high yield (Robertson et al. 1996), which hydrolyses to U and C, providing a mechanism for the accumulation of pyrimidines in the prebiotic environment. A eutectic reaction producing the biological pyrimidines has also been demonstrated (Cleaves et al. 2006). U (albeit in low yields) and its biosynthetic precursor orotic acid were also identified in the hydrolysis products of HCN polymer (Ferris et al. 1978; Voet and Schwartz 1982), and U has been identified in CCs (Stoks and Schwartz 1979).

The reaction of U with formaldehyde and formate gives thymine (T) in good yield (Choughuley et al. 1977). T is also synthesized from the UV-catalyzed dehydrogenation of dihydrothymine, produced from the reaction of β-aminoisobutyric acid with urea (Schwartz and Chittenden 1977).

Sugars

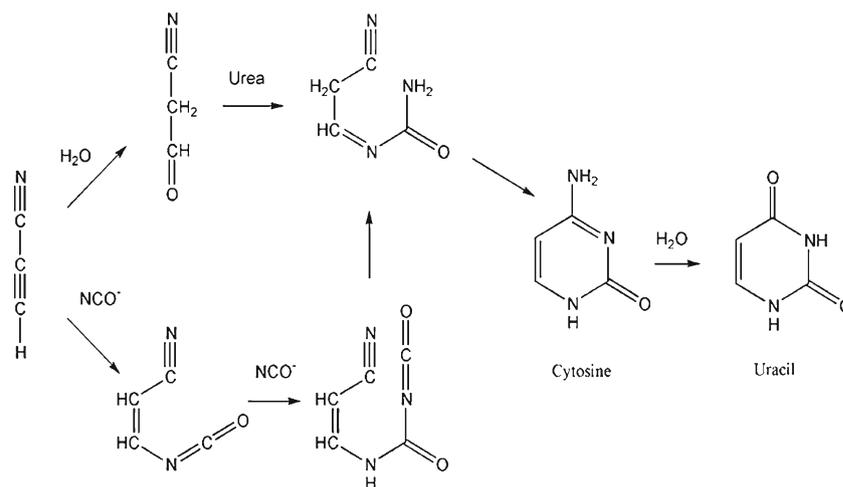
Many biological sugars have the empirical formula (CH₂O)_n, a point underscored by Butlerov's 1861 discovery of the formose reaction, which showed that a diverse assortment of sugars can be formed by the reaction of HCHO under basic conditions (Butlerow 1861). The Butlerov synthesis is complex and incompletely understood. It depends on the presence of catalysts, with Ca(OH)₂ or CaCO₃ being the most completely investigated. In the absence of base catalysts, little or no sugar is obtained. Clays such as kaolin catalyze the formation of sugars, including ribose, in small yields from dilute (0.01 molar) HCHO solutions (Gabel and Ponnampuruma 1967; Reid and Orgel 1967; Schwartz and Degraaf 1993).

The Butlerov synthesis is autocatalytic and proceeds through glycoaldehyde, glyceraldehyde, and dihydroxyacetone. The reaction is also catalyzed by glycolaldehyde, acetaldehyde, and various other organic catalysts (Matsumoto et al. 1984).

The reaction may proceed as shown in Fig. 8. The reaction tends to stop when the formaldehyde has been consumed and ends with the production of C4–C7 sugars that can form cyclic acetals and ketals.

The reaction produces all of the epimers and isomers of the small C2–C6 sugars, some of the C7 ones, and various dendroaldoses and dendroketoses, as well as polyols such as

Fig. 7 Possible mechanisms for the prebiotic synthesis of pyrimidines from the reaction of HCCCN or CAA with urea concentrated in drying lagoon models



glycerol and pentaerythritol, and is generally not particularly selective, although methods of overcoming this have been investigated. Of special interest is the observation that borate can skew the product mixture in favor of certain sugars, including ribose (Prieur 2001; Ricardo et al. 2004). Inclusion of acetaldehyde in the reaction produces deoxyribose (Oró 1965).

Problems with the formose reaction as a source of sugars on the primitive Earth have been noted. One is the complexity of the product mixture. More than 40 different sugars were identified in one reaction mixture (Decker et al. 1982). Another problem is that the conditions of synthesis are also

conducive to the degradation of sugars (Reid and Orgel 1967). Sugars undergo various reactions on short geological time scales that are seemingly prohibitive to their accumulation in the environment. At pH seven, the half-life for ribose decomposition is 73 minutes at 100 degrees Celsius and 44 years at zero degree Celsius (Larralde et al. 1995). Most other sugars are similarly labile.

Additionally, when aqueous solutions of HCN and HCHO are mixed, the major product is glycolonitrile (Schlesinger and Miller 1973), which could preclude the formation of sugars and purines in the same location (Arrhenius et al. 1994). Nevertheless, both sugar derivatives and nucleic acid bases

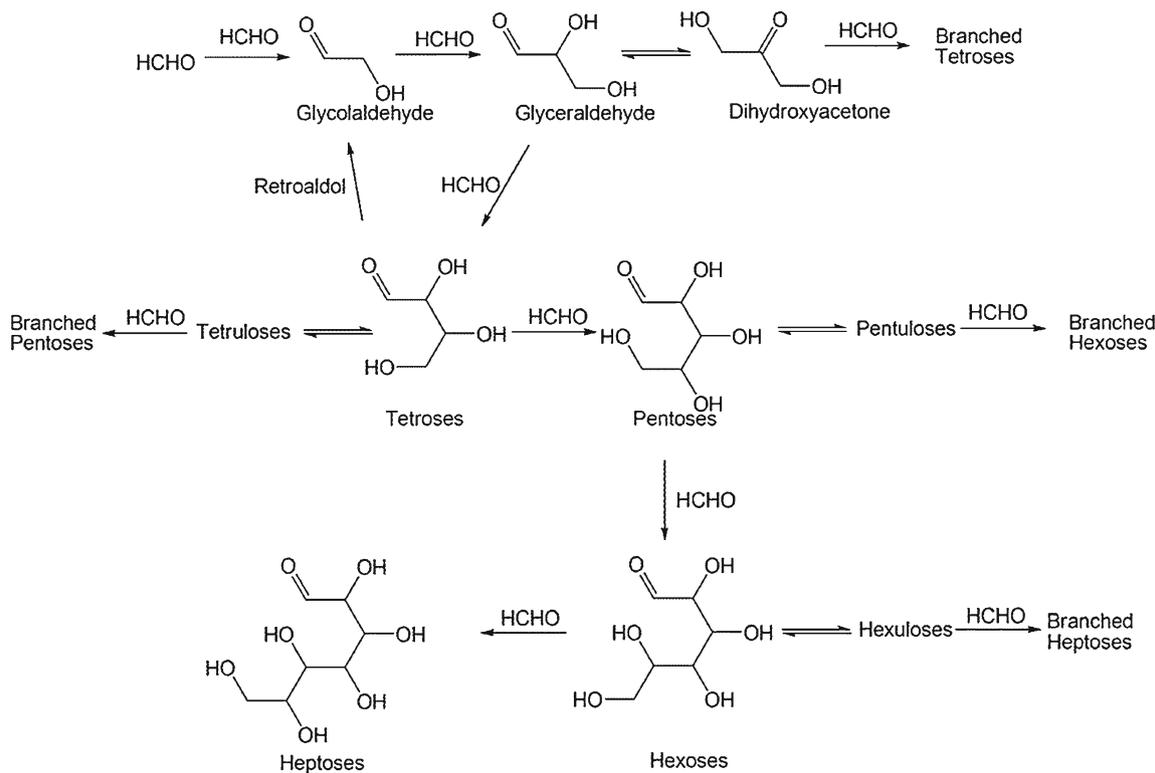


Fig. 8 A simplified scheme of the formose reaction

have been found in the Murchison meteorite (Cooper et al. 2001; Callahan et al. 2011), and it seems likely that the chemistry which produced the compounds found in Murchison meteorite was from aqueous reactions of simple species such as HCN and HCHO. This suggests that the synthesis of sugars, amino acids, and purines from HCHO and HCN may take place under certain conditions.

Nucleosides, Nucleotides

The earliest attempts to produce nucleosides prebiotically involved simply heating purines or pyrimidines in the dry state with ribose (Fuller et al. 1972). Using hypoxanthine and a mixture of salts reminiscent of those found in seawater, up to eight % of β -D-inosine was formed, along with the α -isomer. Adenine and guanine gave lower yields, and in both cases, a mixture of α - and β -isomers was obtained (Fuller et al. 1972). Direct heating of ribose and U or C has thus far failed to produce uridine or cytidine. Pyrimidine nucleoside syntheses have been demonstrated which start from ribose, cyanamide, and cyanoacetylene; however, α -D-cytidine is the major product (Sanchez and Orgel 1970). This can be photo-anomerized to β -D-cytidine in low yield. Sutherland and coworkers (Ingar et al. 2003) demonstrated that cytidine-3'-phosphate can be prepared from arabinose-3-phosphate, cyanamide, and HCCCN in a one-pot reaction. The prebiotic source of arabinose-3-phosphate is unclear; nevertheless, it remains possible that more creative methods of prebiotic pyrimidine nucleoside synthesis can be found.

More recently, reactions using more complex reagents added in precise orders, with the products isolated and carried through to the next step, have been shown to produce significant yields of ribotides (Powner et al. 2007). These reactions give good yields of the pyrimidine ribotides (Powner and Sutherland 2008), and a related series of reactions could produce the purine ribotides (Powner et al. 2010). It is not clear that these syntheses solve previously raised objections to the plausibility of prebiotic nucleoside and nucleotide synthesis (Shapiro 1984).

Prebiotic phosphorylation of nucleosides has also been demonstrated, but again with caveats. Small amounts of condensed phosphates are emitted in volcanic fumaroles (Yamagata et al. 1991), and heating orthophosphate at relatively low temperatures in the presence of ammonia results in a high yield of condensed phosphates (Osterberg and Orgel 1972). Trimetaphosphate (TMP) has been shown to be an active phosphorylating agent for various molecules including amino acids and nucleosides (Schwartz 1969; Rabinowitz and Hampai 1978; Yamagata et al. 1991). However, it has also been suggested that condensed phosphates are not likely to be prebiotically abundant materials (Keefe and Miller 1995).

Early attempts to produce nucleotides using organic condensing reagents such as H_2CN , NCO^- , or dicyanamide

(Lohrmann and Orgel 1968) were generally inefficient due to the competition of the alcohol groups of the nucleosides with water in an aqueous environment. Nucleosides can be phosphorylated with acidic phosphates such as NaH_2PO_4 when dry heated (Beck et al. 1967). These reactions are catalyzed by urea and other amides, particularly if NH_3 is included in the reaction. Nucleosides can also be phosphorylated in high yield by heating ammonium phosphate with urea at moderate temperatures, as might occur in a drying lagoon (Lohrmann and Orgel 1971). Heating uridine with urea and ammonium phosphate gave yields of nucleotides as high as 70%. In the case of purine nucleotides, however, there is also considerable glycosidic cleavage due to the acidic microenvironment created. Thus, another problem with the “prebiotic” RNA world is that the synthesis of purine nucleosides is somewhat robust, but nucleotide formation may be difficult, while nucleotide formation from pyrimidine nucleosides is robust, but nucleoside formation may be difficult.

Common calcium phosphate minerals such as apatite are themselves reasonable phosphorylating reagents. Yields as high as 20% of nucleotides were achieved by heating nucleosides with apatite, urea, and ammonium phosphate (Lohrmann and Orgel 1971). Heating ammonium phosphates with urea leads to a mixture of high molecular weight polyphosphates (Osterberg and Orgel 1972). Although these are not especially good phosphorylating reagents under prebiotic conditions, they may degrade, especially in the presence of divalent cations at high temperatures, to cyclic phosphates such as TMP.

The difficulties with prebiotic ribose synthesis and nucleoside formation have led some to speculate that perhaps a simpler genetic molecule with a more robust prebiotic synthesis preceded RNA (Joyce et al. 1987). Substituting sugars besides ribose has been proposed (Eschenmoser 2004). Oligomers of some of these also form stable base-paired structures with both RNA/DNA and themselves, opening the possibility of genetic takeover from a precursor polymer to RNA/DNA. Such molecules may suffer from the same drawbacks as RNA with respect to prebiotic chemistry, such as the difficulty of selective sugar synthesis, sugar instability, and the difficulty of nucleoside formation. It has been demonstrated based on the suggestion of Joyce et al. (1987) and proposed chemistry (Nelsestuen 1980; Tohidi and Orgel 1989) that backbones based on acyclic nucleoside analogs may be more easily obtained under reasonable prebiotic conditions, for example by the reaction of nucleobases with acrolein obtained from mixed formose reactions (Cleaves 2002).

More exotic alternatives to nucleosides have been proposed, for example peptide nucleic acid (PNA) analogs (Nielsen et al. 1994). Miller and coworkers were able to demonstrate the prebiotic synthesis of the components of PNA under the same chemical conditions required for the synthesis of the purines or pyrimidines (Nelson et al. 2000).

The vast majority of possible alternative structures have not been investigated with respect to prebiotic plausibility.

Environmental Considerations

Whether in meteorites or on Earth, prebiotic chemistry may have occurred largely in an aqueous environment, as water is a ubiquitous component of the solar system and the Earth's surface. Among the variables of the local environment which could affect the way this chemistry occurs are pH, temperature, inorganic compounds such as metals, mineral surfaces, the impact of sunlight, etc. The potential role of mineral surfaces on prebiotic chemistry is an especially complex and under-explored aspect of this chemistry. Although we do not presently know which compounds were essential for the origin of life, it seems possible to preclude certain environments if even fairly simple organic compounds were involved (Cleaves and Chalmers 2004).

Stability of Biomolecules at High Temperatures

High temperatures cause reactions to occur more quickly, they are also destructive to most bioorganic compounds. Although some progress has been made in synthesizing small molecules under conditions simulating hydrothermal vents, most biological molecules have hydrolytic half-lives on the order of minutes to seconds at the temperatures associated with on-axis hydrothermal vents and are still rather unstable at the lower temperatures of off-axis vent environments. Ribose and other sugars are extremely thermolabile (Larralde et al. 1995), but pyrimidines and purines, and many amino acids, are nearly as labile. At 100 degrees Celsius, the half-life ($t_{1/2}$) for deamination of cytosine is 21 days, and 204 days for adenine (Levy and Miller 1998). Some amino acids are more stable: alanine has a $t_{1/2}$ for decarboxylation of $\sim 1.9 \times 10^4$ years at 100 degrees Celsius, but serine decarboxylates to ethanolamine with a $t_{1/2}$ of 320 days (Vallentyne 1964). White (1984) measured the decomposition of various compounds at 250 degrees Celsius and pH seven and found $t_{1/2}$ values for amino acids from 7.5 seconds to 278 minutes, peptide bonds from less than one minute to 11.8 minutes, glycoside cleavage in nucleosides from less than one second to 1.3 minutes, decomposition of nucleobases from 15 to 57 minutes, and phosphate esters from 2.3 to 420 minutes. The half-lives for polymers would be even shorter as there are so many more potential breakage points.

Minerals

There are approximately 4,400 known naturally occurring minerals on Earth today (Hazen et al. 2008). This number

was likely smaller on the early Earth, as many minerals are produced by oxidation with environmental O_2 , biological deposition, or the vast amounts of time which have passed since the Earth formed.

Minerals may have complex effects on prebiotic organic synthesis (Lahav and Chang 1976) by concentrating reactants and by lowering activation barriers to bring compounds into more rapid equilibrium (Marshall-Bowman et al. 2010).

Shallow Pools

Since Darwin's time, it has been supposed that life may have originated in shallow tidal pools. Possible advantages of such environments include evaporative concentration, constrained diffusion mineral catalysis, and the penetration of sunlight into such pools. Additionally, if the Earth's surface was colder at this time, or at least as clement as at present, then it is possible that periodic freezing events could have significantly concentrated compounds.

Hydrothermal Vents

The discovery of hydrothermal vents at mid-ocean ridges and the appreciation of their significance in the element balance of the hydrosphere were a major discovery in oceanography (Corliss et al. 1979). Since the process of hydrothermal circulation probably began early in Earth's history, it is likely that vents were also present. Large amounts of ocean water now pass through the vents, with the whole ocean going through them approximately every ten million years (Edmond et al. 1982). This flow was probably greater during the early history of the Earth, since the heat flow from the planet's interior was greater. The topic has received a great deal of attention, partly because of uncertainty regarding the oxidation state of the early atmosphere.

There are various types of hydrothermal environments on the modern Earth, including subaerial hot springs and submarine hydrothermal vents. In the latter, temperatures range from \sim four to 350 degrees Celsius, with pH ranging from zero to 11 (Martin et al. 2008). Various minerals precipitate as the heated vent water enters the surrounding ocean water, leading to the formation of baroque rock formations. It has been speculated that the pores in such minerals may have served to concentrate organic species via thermophoresis (Baaske et al. 2007).

Following the vents' discovery, a hypothesis suggesting a hydrothermal emergence of life was published (Corliss et al. 1981), which suggested that amino acids and other organic compounds could be produced during passage of the effluent through the temperature gradient from \sim 350 degrees Celsius to \sim zero degree Celsius, roughly the temperature of modern ambient ocean waters.

Polymerization of the organic compounds thus formed, followed by their self-organization, was also proposed in this environment.

At first glance, submarine hydrothermal springs appear to be ideally suited for organic synthesis, given the geological plausibility of a hot early Earth. Vents exist along the active tectonic areas of the Earth, and in at least some of them, potentially catalytic minerals interact with an aqueous reducing environment rich in H_2 , H_2S , CO , CO_2 , CH_4 , and NH_3 . Unfortunately, it is difficult to corroborate these speculations with the composition of the effluents of modern vents, as much of the organic material released from modern sources is simply environmentally processed biological material.

Modern hydrothermal vent fluids do contain some organic compounds, though it remains unclear what percentage are derived from organisms living in the vents or biologically derived matter entrained into the vent fluids and reworked, and what percentage is derived from truly abiotic processes. For example, for amino acids, the bulk of the evidence available supports a biological origin (Bassez et al. 2009). Presently, the amount and type of organic matter found in hydrothermal vent environments unequivocally thought to be of abiotic origin are limited to a few parts per million of small hydrocarbons such as CH_4 and ethane (McCollom et al. 2010).

One of the most articulate autotrophic vent origin-of-life hypotheses stems from the work of Wächtershäuser (1988, 1992), who argued that life began with the appearance of an autocatalytic, two-dimensional chemolithotrophic metabolic system based on the formation of the insoluble mineral pyrite (FeS_2). The FeS/H_2S combination is a strong reducing agent and has been shown to reduce some organic compounds under mild conditions. Wächtershäuser's scenario fits well with the environmental conditions found at deep-sea hydrothermal vents, where H_2S , CO_2 , and CO are abundant; however, the FeS/H_2S system does not reduce CO_2 to amino acids, purines, or pyrimidines, although there is more than enough free energy to do so (Keefe et al. 1995). Pyrite formation can produce molecular hydrogen and reduce nitrate to NH_3 , and $HCCH$ to H_2CCH_2 (Maden 1995). More recent experiments have shown that the activation of amino acids with carbon monoxide and $(Ni, Fe)S$ can lead to peptide bond formation (Huber and Wächtershäuser 1998), though the degree to which these experiments mimic geochemical environments is a subject of debate.

In general, organic compounds are decomposed rather than created at hydrothermal vent temperatures, although temperature gradients exist. Sowerby and coworkers have shown (Sowerby et al. 2001) that concentration on mineral surfaces would tend to concentrate organics created at hydrothermal vents in cooler zones. The presence of reduced metals at high temperatures could facilitate Fischer–Tropsch-type (FTT)

syntheses. FTT catalysts are poisoned by water and sulfide, though some of the likely catalysts such as magnetite may be immune to such poisoning (Holm and Andersson 1998).

Submarine hydrothermal vents do not seem to presently synthesize organic compounds more complex than simple hydrocarbons such as CH_4 and ethane (McCollom et al. 2010). More likely, they decompose them over short time spans ranging from seconds to a few hours. The origin of life in hydrothermal vents thus may be problematic.

This does not imply that hydrothermal vents were a negligible factor on the primitive Earth. If mineral assemblages were sufficiently reducing, vents may have been a source of atmospheric NH_3 , CO , CH_4 , and H_2 . The concentrations of biomolecules in primitive Earth environments would have been governed by the rates of production and the rates of destruction. Submarine hydrothermal vents would have likely been more important for the destruction of organic compounds, fixing the upper limit for their concentrations in the primitive oceans.

Prebiotic Chemistry Beyond Earth

Of the eight accepted planets in our solar system and their moons, several appear compatible with the synthesis of organic compounds, and several are known to contain them. Fewer appear to be compatible with the existence of liquid water or the more complicated evolution of these compounds. For example, the extreme temperatures of Venus' or Mercury's sunlit side are likely too hostile for the synthesis of complex organics. The immediate sub-surface of Mars appears to harbor both liquid and solid water (Rennó et al. 2009), and it is widely believed that liquid water once flowed on Mars' surface. Meteorites likely impacted Mars' surface at this time, and it is reasonable to expect that some of these were CCs. The nature of Mars' early atmosphere remains unknown, but Mars and Earth may have been similar early in their history.

Proceeding outward from the Sun, complex organic chemistry likely occurs in the atmospheres of Saturn and Jupiter. The conditions on the solid surfaces of these planets are thought to be too hostile for more complex organic chemistry. Nevertheless, the presence of various organic species has been confirmed in their atmospheres (Lodders 2010).

Outer-Planet Moons

A number of outer-planet moons have intriguing environments which appear to foster prebiotic chemistry and could conceivably be capable of sustaining biology. For example, Saturn's moon Titan is now known to harbor a rich organic chemistry (Waite et al. 2007). Jupiter's moon Europa is covered with a thick ice layer which may harbor a watery

ocean several kilometers thick (Manga and Wang 2007). If this ocean does exist, its organic content remains unknown.

Extrasolar Planets

For centuries, the existence of planets beyond our solar system has been the subject of intense philosophical and scientific speculation (Urey et al. 1963; Dick 1999). The first extrasolar planet was confirmed in 2003 (Hatzes et al. 2003). Recent astronomical observations have since yielded hundreds of planets orbiting other suns (<http://www.exoplanet.eu/>; <http://nsted.ipac.caltech.edu/>). Methods to date have favored the detection of large planets with orbits very close to their parent suns, but soon, advances will allow detection of smaller planets with orbits compatible with the existence of liquid surface water. Our solar system contains three rocky Earth-like planets, two of which are *almost* in the habitable zone.

There is no solid reason to expect that our solar system is anomalous. Given the billions of Sun-like stars in any galaxy, there may be many rocky Earth-like planets in stellar habitable zones in the universe (Laughlin 2010), some of which may have undergone similar periods of evolution compatible with the origin and evolution of life. The detection of such planets and possible signatures of alien biochemistry may not be far off, assuming that the origin of life is facile. The detection of even one such planet would strongly reinforce the idea that it is.

Unknowns and the Future

A basic tenet of the heterotrophic theory of the origin of life is that the origin of the first living systems depended on environmentally supplied organic molecules. As summarized here, there has been no shortage of discussion as to how the formation of these molecules occurred. Organic compounds may have accumulated on the primitive Earth via numerous mechanisms including contributions from endogenous atmospheric synthesis, deep-sea hydrothermal vent synthesis, and exogenous delivery. Though this raises the issue of the relative significance of the various sources, it also recognizes the wide variety of potential sources of organic compounds.

Given adequate expertise and experimental conditions, it is possible to synthesize many organic molecules in the laboratory under simulated prebiotic conditions. The fact that a number of molecular components of contemporary cells can be formed in the laboratory does not necessarily mean that they were essential for the origin of life, or that they were likely prebiotically available. The Earths' primordial soup may have been a complex mixture, but it likely did not include all of the compounds found today in even the simplest cells.

There are many mechanisms by which biochemical monomers can be prebiotically synthesized. Not all prebiotic pathways are equally efficient, but the wide range of conditions under which organic compounds can be synthesized demonstrates that prebiotic syntheses of the building blocks of life are robust, i.e., they can occur in a wide variety of environmental settings. Although our ideas regarding the prebiotic synthesis of organic compounds are based largely on model experiments, the robustness of this type of chemistry is supported by the occurrence of many of these compounds in CCs. It is therefore plausible that a similar synthesis took place on the primitive Earth.

We do not yet have a coherent model for how life arose on Earth or how it might arise on other planets. We do, however, have evidence that many organic compounds, including many of those crucial to the functioning of modern organisms, are easily produced via prebiotic chemistry and are likely widespread throughout the universe. We also know that in some cases, some of these spontaneously self-organize.

It is impossible to predict how soon the problems prebiotic chemistry addresses will be solved. It seems likely that in the near future, chemists will assemble living organisms *de novo*. As of this writing, an entire synthetic genome was constructed (Gibson et al. 2008), a ribozyme capable of assembling 80% of its sequence was developed in the laboratory (Wochner et al. 2011), and a self-sustained and autocatalytic set of self-propagating RNA molecules had been culled from a laboratory experiment (Lincoln and Joyce 2009).

What remains more problematic is a comprehensive understanding of how such systems may have arisen from environmentally plausible reactions on the primitive Earth. Despite considerable progress, much work remains, and undoubtedly, many surprises are in store.

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